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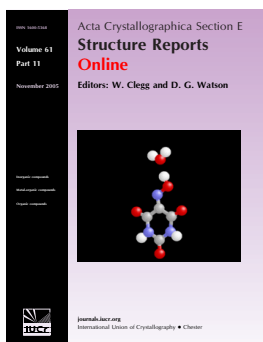
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Scheelite-type sodium neodymium(III) *ortho*-oxidomolybdate(VI), $\text{NaNd}[\text{MoO}_4]_2$

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Scheelite-type sodium neodymium(III) ortho-oxidomolybdate(VI), NaNd[MoO₄]₂

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Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{Mo}-\text{O}) = 0.002$ Å; disorder in main residue; R factor = 0.020; wR factor = 0.044; data-to-parameter ratio = 13.1.

Scheelite-type NaNd[MoO₄]₂ contains one crystallographic position (site symmetry $\bar{4}$) for the large cations, which is mixed-occupied by Na⁺ and Nd³⁺ cations in a 1:1 molar ratio. Thus, both are surrounded by eight O atoms in the shape of a trigonal dodecahedron. Furthermore, the structure consists of crystallographically unique [MoO₄]²⁻ units (site symmetry $\bar{4}$) surrounded by eight sodium and neodymium cations, which are all vertex-attached. The polyhedra around the Na⁺/Nd³⁺ cations are connected to four others *via* common edges, building up a three-dimensional network in whose tetrahedral voids of O atoms the Mo⁶⁺ cations reside.

Related literature

For isotypic NaLn[MoO₄]₂ structures, see: Stevens *et al.* (1991) and Teller (1992) for $Ln = \text{La}$; Teller (1992) for $Ln = \text{Ce}$; Zhao *et al.* (2010) for $Ln = \text{Er}$. For interpenetrating diamond-like networks, see: Schustereit *et al.* (2011). These were also observed in NaTl, see: Zintl & Dullenkopf (1932).

Experimental

Crystal data

NaNd[MoO₄]₂ $M_r = 487.11$

Tetragonal, $I4_1/a$
 $a = 5.2871$ (3) Å
 $c = 11.5729$ (7) Å
 $V = 323.50$ (3) Å³
 $Z = 2$

Mo $K\alpha$ radiation
 $\mu = 11.79$ mm⁻¹
 $T = 293$ K
 $0.11 \times 0.09 \times 0.07$ mm

Data collection

Nonius KappaCCD diffractometer
 Absorption correction: numerical
 (*X-SHAPE*; Stoe & Cie, 1995)
 $T_{\min} = 0.273$, $T_{\max} = 0.434$

1026 measured reflections
 196 independent reflections
 135 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.058$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.020$
 $wR(F^2) = 0.044$
 $S = 0.99$
 196 reflections

15 parameters
 $\Delta\rho_{\text{max}} = 0.45$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.42$ e Å⁻³

Data collection: *COLLECT* (Nonius, 1998); cell refinement: *SCALEPACK* (Otwinowski & Minor, 1997); data reduction: *SCALEPACK* and *DENZO* (Otwinowski & Minor, 1997); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *DIAMOND* (Brandenburg, 2006); software used to prepare material for publication: *SHELXL97*.

This work was supported by the State of Baden-Württemberg (Stuttgart) and the Deutsche Forschungsgemeinschaft (DFG; Frankfurt/Main) within the funding programme Open Access Publishing.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HP2019).

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supplementary materials

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Scheelite-type sodium neodymium(III) *ortho*-oxidomolybdate(VI), NaNd[MoO₄]₂

T. Schleid and I. Hartenbach

Comment

Scheelite-type sodium lanthanide *ortho*-oxomolybdates of the formula NaLn[MoO₄]₂ are already known for Ln = La, Ce, and Er (see related literature). The structure features Na⁺ and Nd³⁺ cations together at the common *Wyckoff* position 4*b*, eightfold coordinated by O²⁻ anions in the shape of trigonal dodecahedra (Fig. 1). The Na : Nd ratio was fixed at a molar ratio of 1 : 1 for maintaining electroneutrality. A similar surrounding is found for the cationic coordination around the crystallographically unique isolated *ortho*-oxomolybdate anions [MoO₄]²⁻ with the Mo⁶⁺ cations at the *Wyckoff* position 4*a* (Fig. 2). The polyhedra around the Na⁺/Nd³⁺ cations are interconnected to four others via common edges building up a three-dimensional network, in whose voids of oxygen the Mo⁶⁺ cations are located (Fig. 3). Both the cations at the sites 4*a* (Na⁺ and Nd³⁺ in a 1:1 molar ratio) and 4*b* (Mo⁶⁺) arrange themselves in two interpenetrating *diamond*-like networks (Schustereit *et al.*, 2011) as in the case of NaTl (Zintl & Dullenkopf, 1932).

Experimental

Pale violet, coarse single crystals of Scheelite-type NaNd[MoO₄]₂ were obtained as by-product in an unsuccessful attempt to synthesize NdF[MoO₄], using a mixture of NdF₃ and Na₂[MoO₄] in a 1:1 molar ratio, which was heated at 850 °C for 7 days in an evacuated, sealed, fused-silica ampoule.

Figures

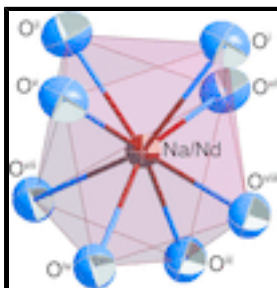


Fig. 1. Trigonal dodecahedral oxygen environment of the Na⁺/Nd³⁺ cations in Scheelite-type NaNd[MoO₄]₂ (ellipsoids are drawn at 90 % probability level, symmetry codes: (i) $y-1/4, -x+3/4, z+3/4$; (ii) $-y+1/4, x-1/4, z+3/4$; (iii) $x-1/2, y, -z+1/2$; (iv) $-x+1/2, -y+1/2, -z+1/2$; (v) $-y+3/4, x-1/4, -z+3/4$; (vi) $y-3/4, -x+3/4, -z+3/4$; (vii) $x-1/2, y-1/2, z+1/2$; (viii) $-x+1/2, -y+1, z+1/2$)

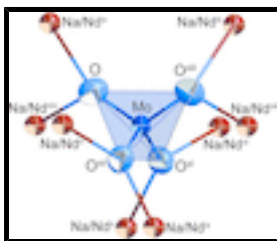


Fig. 2. Cationic surrounding of the isolated *ortho*-oxomolybdate(VI) tetrahedra [MoO₄]²⁻ in Scheelite-type NaNd[MoO₄]₂ (ellipsoids are drawn at 90 % probability level, symmetry codes for O: x, y, z ; (xi) $-y+1/4, x+1/4, -z+1/4$; (xii) $y-1/4, -x+1/4, z+1/4$; (xiii) $-x, -y+1/2, z$; symmetry codes for Na/Nd: (iii) $x-1/2, y, -z+1/2$; (iv) $-x+1/2, -y+1/2, -z+1/2$; (v) $-y+3/4, x-1/4, -z+3/4$; (vi) $y-3/4, -x+3/4, -z+3/4$; (vii) $x-1/2, y-1/2, z+1/2$; (ix) $-x, -y, -z+1$; (x) $-x, -y+1, -z+1$; (xiv) $x+1/2, y+1/2, z-1/2$)

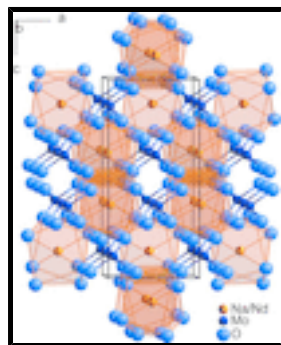


Fig. 3. View at the crystal structure of Scheelite-type $\text{NaNd}[\text{MoO}_4]_2$ along $[010]$ (slightly rotated) with special emphasis on the edge-connected oxygen polyhedra around the Na^+ and Nd^{3+} cations, respectively.

sodium neodymium(III) *ortho*-oxidomolybdate(vi)

Crystal data

| | |
|--------------------------------|---|
| $\text{NaNd}[\text{MoO}_4]_2$ | $D_x = 5.001 \text{ Mg m}^{-3}$ |
| $M_r = 487.11$ | Mo $K\alpha$ radiation, $\lambda = 0.71069 \text{ \AA}$ |
| Tetragonal, $I4_1/a$ | Cell parameters from 2325 reflections |
| Hall symbol: $-I\ 4ad$ | $\theta = 1.0\text{--}28.3^\circ$ |
| $a = 5.2871 (3) \text{ \AA}$ | $\mu = 11.79 \text{ mm}^{-1}$ |
| $c = 11.5729 (7) \text{ \AA}$ | $T = 293 \text{ K}$ |
| $V = 323.50 (3) \text{ \AA}^3$ | Coarse transparent, pale violet |
| $Z = 2$ | $0.11 \times 0.09 \times 0.07 \text{ mm}$ |
| $F(000) = 438$ | |

Data collection

| | |
|---|--|
| Nonius KappaCCD diffractometer | 196 independent reflections |
| Radiation source: fine-focus sealed tube graphite | 135 reflections with $I > 2\sigma(I)$ |
| ω and ϕ scans | $R_{\text{int}} = 0.058$ |
| Absorption correction: numerical (<i>X-SHAPE</i> ; Stoe & Cie, 1995) | $\theta_{\text{max}} = 28.2^\circ$, $\theta_{\text{min}} = 4.2^\circ$ |
| $T_{\text{min}} = 0.273$, $T_{\text{max}} = 0.434$ | $h = -6 \rightarrow 6$ |
| 1026 measured reflections | $k = -6 \rightarrow 6$ |
| | $l = -15 \rightarrow 15$ |

Refinement

| | |
|---------------------------------|--|
| Refinement on F^2 | Primary atom site location: structure-invariant direct methods |
| Least-squares matrix: full | Secondary atom site location: difference Fourier map |
| $R[F^2 > 2\sigma(F^2)] = 0.020$ | $w = 1/[\sigma^2(F_o^2) + (0.0102P)^2]$ |
| $wR(F^2) = 0.044$ | where $P = (F_o^2 + 2F_c^2)/3$ |
| $S = 0.99$ | $(\Delta/\sigma)_{\text{max}} < 0.001$ |
| 196 reflections | $\Delta\rho_{\text{max}} = 0.45 \text{ e \AA}^{-3}$ |
| | $\Delta\rho_{\text{min}} = -0.42 \text{ e \AA}^{-3}$ |

15 parameters

Extinction correction: *SHELXL97* (Sheldrick, 2008),

$$F_c^* = kFc[1+0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$$

0 restraints

Extinction coefficient: 0.0092 (8)

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R-factor wR and goodness of fit S are based on F^2 , conventional R-factors R are based on F, with F set to zero for negative F^2 . The threshold expression of $F^2 > 2\text{sigma}(F^2)$ is used only for calculating R-factors(gt) etc. and is not relevant to the choice of reflections for refinement. R-factors based on F^2 are statistically about twice as large as those based on F, and R- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

| | x | y | z | $U_{\text{iso}}^*/U_{\text{eq}}$ | Occ. (<1) |
|----|------------|------------|--------------|----------------------------------|-----------|
| Na | 0.0000 | 0.2500 | 0.6250 | 0.0125 (2) | 0.50 |
| Nd | 0.0000 | 0.2500 | 0.6250 | 0.0125 (2) | 0.50 |
| Mo | 0.0000 | 0.2500 | 0.1250 | 0.0131 (2) | |
| O | 0.2406 (3) | 0.3949 (3) | 0.04140 (14) | 0.0255 (5) | |

Atomic displacement parameters (\AA^2)

| | U^{11} | U^{22} | U^{33} | U^{12} | U^{13} | U^{23} |
|----|-------------|-------------|------------|-------------|------------|-------------|
| Na | 0.0130 (3) | 0.0130 (3) | 0.0113 (3) | 0.000 | 0.000 | 0.000 |
| Nd | 0.0130 (3) | 0.0130 (3) | 0.0113 (3) | 0.000 | 0.000 | 0.000 |
| Mo | 0.0122 (3) | 0.0122 (3) | 0.0151 (3) | 0.000 | 0.000 | 0.000 |
| O | 0.0298 (10) | 0.0227 (12) | 0.0239 (9) | -0.0002 (8) | 0.0022 (8) | -0.0008 (9) |

Geometric parameters (\AA , $^\circ$)

| | | | |
|-------------------------------------|--------------|---------------------------------------|--------------|
| Na—O ⁱ | 2.4851 (15) | Na—Nd ^x | 3.91907 (16) |
| Na—O ⁱⁱ | 2.4851 (15) | Na—Na ^x | 3.91907 (16) |
| Na—O ⁱⁱⁱ | 2.4851 (15) | Mo—O ^{xi} | 1.7725 (16) |
| Na—O ^{iv} | 2.4851 (15) | Mo—O | 1.7725 (17) |
| Na—O ^v | 2.5182 (18) | Mo—O ^{xii} | 1.7725 (16) |
| Na—O ^{vi} | 2.5182 (18) | Mo—O ^{xiii} | 1.7725 (16) |
| Na—O ^{vii} | 2.5182 (18) | O—Nd ⁱⁱⁱ | 2.4851 (15) |
| Na—O ^{viii} | 2.5182 (18) | O—Na ⁱⁱⁱ | 2.4851 (15) |
| Na—Na ^{ix} | 3.91907 (17) | O—Na ^{xiv} | 2.5182 (18) |
| Na—Nd ^{ix} | 3.91907 (17) | O—Nd ^{xiv} | 2.5182 (18) |
| O ⁱ —Na—O ⁱⁱ | 126.90 (5) | O ^{vi} —Na—Nd ^{ix} | 102.66 (4) |
| O ⁱ —Na—O ⁱⁱⁱ | 126.90 (5) | O ^{vii} —Na—Nd ^{ix} | 85.20 (3) |

supplementary materials

| | | | |
|--|------------|--|-------------|
| O ⁱⁱ —Na—O ⁱⁱⁱ | 78.41 (7) | O ^{viii} —Na—Nd ^{ix} | 130.62 (4) |
| O ⁱ —Na—O ^{iv} | 78.41 (7) | Na ^{ix} —Na—Nd ^{ix} | 0.0 |
| O ⁱⁱ —Na—O ^{iv} | 126.90 (5) | O ⁱ —Na—Nd ^x | 38.74 (4) |
| O ⁱⁱⁱ —Na—O ^{iv} | 126.90 (5) | O ⁱⁱ —Na—Nd ^x | 160.78 (4) |
| O ⁱ —Na—O ^v | 151.28 (7) | O ⁱⁱⁱ —Na—Nd ^x | 101.53 (4) |
| O ⁱⁱ —Na—O ^v | 68.38 (4) | O ^{iv} —Na—Nd ^x | 68.65 (4) |
| O ⁱⁱⁱ —Na—O ^v | 76.88 (6) | O ^v —Na—Nd ^x | 130.62 (4) |
| O ^{iv} —Na—O ^v | 73.65 (3) | O ^{vi} —Na—Nd ^x | 85.20 (3) |
| O ⁱ —Na—O ^{vi} | 73.65 (3) | O ^{vii} —Na—Nd ^x | 38.14 (3) |
| O ⁱⁱ —Na—O ^{vi} | 76.88 (6) | O ^{viii} —Na—Nd ^x | 102.66 (4) |
| O ⁱⁱⁱ —Na—O ^{vi} | 68.38 (4) | Na ^{ix} —Na—Nd ^x | 123.025 (3) |
| O ^{iv} —Na—O ^{vi} | 151.28 (7) | Nd ^{ix} —Na—Nd ^x | 123.025 (3) |
| O ^v —Na—O ^{vi} | 134.81 (8) | O ⁱ —Na—Na ^x | 38.74 (4) |
| O ⁱ —Na—O ^{vii} | 76.88 (6) | O ⁱⁱ —Na—Na ^x | 160.78 (4) |
| O ⁱⁱ —Na—O ^{vii} | 151.28 (7) | O ⁱⁱⁱ —Na—Na ^x | 101.53 (4) |
| O ⁱⁱⁱ —Na—O ^{vii} | 73.65 (3) | O ^{iv} —Na—Na ^x | 68.65 (4) |
| O ^{iv} —Na—O ^{vii} | 68.38 (4) | O ^v —Na—Na ^x | 130.62 (4) |
| O ^v —Na—O ^{vii} | 98.49 (3) | O ^{vi} —Na—Na ^x | 85.20 (3) |
| O ^{vi} —Na—O ^{vii} | 98.49 (3) | O ^{vii} —Na—Na ^x | 38.14 (3) |
| O ⁱ —Na—O ^{viii} | 68.38 (4) | O ^{viii} —Na—Na ^x | 102.66 (4) |
| O ⁱⁱ —Na—O ^{viii} | 73.65 (3) | Na ^{ix} —Na—Na ^x | 123.025 (3) |
| O ⁱⁱⁱ —Na—O ^{viii} | 151.28 (7) | Nd ^{ix} —Na—Na ^x | 123.025 (3) |
| O ^{iv} —Na—O ^{viii} | 76.88 (6) | Nd ^x —Na—Na ^x | 0.0 |
| O ^v —Na—O ^{viii} | 98.49 (3) | O ^{xi} —Mo—O | 113.83 (11) |
| O ^{vi} —Na—O ^{viii} | 98.49 (3) | O ^{xi} —Mo—O ^{xii} | 107.34 (5) |
| O ^{vii} —Na—O ^{viii} | 134.81 (8) | O—Mo—O ^{xii} | 107.34 (5) |
| O ⁱ —Na—Na ^{ix} | 160.78 (4) | O ^{xi} —Mo—O ^{xiii} | 107.34 (5) |
| O ⁱⁱ —Na—Na ^{ix} | 68.65 (4) | O—Mo—O ^{xiii} | 107.34 (5) |
| O ⁱⁱⁱ —Na—Na ^{ix} | 38.74 (4) | O ^{xii} —Mo—O ^{xiii} | 113.83 (11) |
| O ^{iv} —Na—Na ^{ix} | 101.53 (4) | Mo—O—Nd ⁱⁱⁱ | 133.30 (9) |
| O ^v —Na—Na ^{ix} | 38.14 (3) | Mo—O—Na ⁱⁱⁱ | 133.30 (9) |
| O ^{vi} —Na—Na ^{ix} | 102.66 (4) | Nd ⁱⁱⁱ —O—Na ⁱⁱⁱ | 0.0 |
| O ^{vii} —Na—Na ^{ix} | 85.20 (3) | Mo—O—Na ^{xiv} | 120.23 (8) |
| O ^{viii} —Na—Na ^{ix} | 130.62 (4) | Nd ⁱⁱⁱ —O—Na ^{xiv} | 103.12 (6) |
| O ⁱ —Na—Nd ^{ix} | 160.78 (4) | Na ⁱⁱⁱ —O—Na ^{xiv} | 103.12 (6) |
| O ⁱⁱ —Na—Nd ^{ix} | 68.65 (4) | Mo—O—Nd ^{xiv} | 120.23 (8) |
| O ⁱⁱⁱ —Na—Nd ^{ix} | 38.74 (4) | Nd ⁱⁱⁱ —O—Nd ^{xiv} | 103.12 (6) |
| O ^{iv} —Na—Nd ^{ix} | 101.53 (4) | Na ⁱⁱⁱ —O—Nd ^{xiv} | 103.12 (6) |
| O ^v —Na—Nd ^{ix} | 38.14 (3) | Na ^{xiv} —O—Nd ^{xiv} | 0.0 |

Symmetry codes: (i) $y-1/4, -x+3/4, z+3/4$; (ii) $x-1/2, y, -z+1/2$; (iii) $-x+1/2, -y+1/2, -z+1/2$; (iv) $-y+1/4, x-1/4, z+3/4$; (v) $x-1/2, y-1/2, z+1/2$; (vi) $-x+1/2, -y+1, z+1/2$; (vii) $-y+3/4, x-1/4, -z+3/4$; (viii) $y-3/4, -x+3/4, -z+3/4$; (ix) $-x, -y, -z+1$; (x) $-x+1/2, -y+1/2, -z+3/2$; (xi) $-x, -y+1/2, z$; (xii) $-y+1/4, x+1/4, -z+1/4$; (xiii) $y-1/4, -x+1/4, -z+1/4$; (xiv) $x+1/2, y+1/2, z-1/2$.

Fig. 1

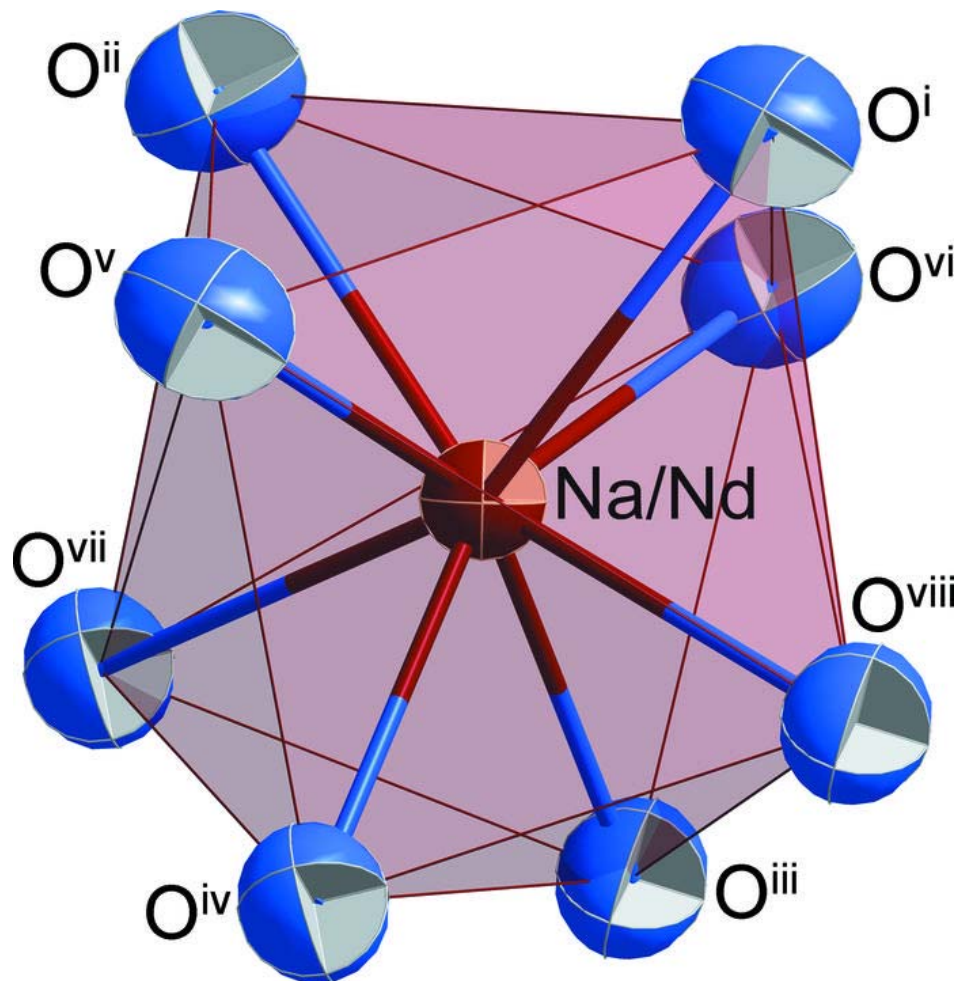


Fig. 2

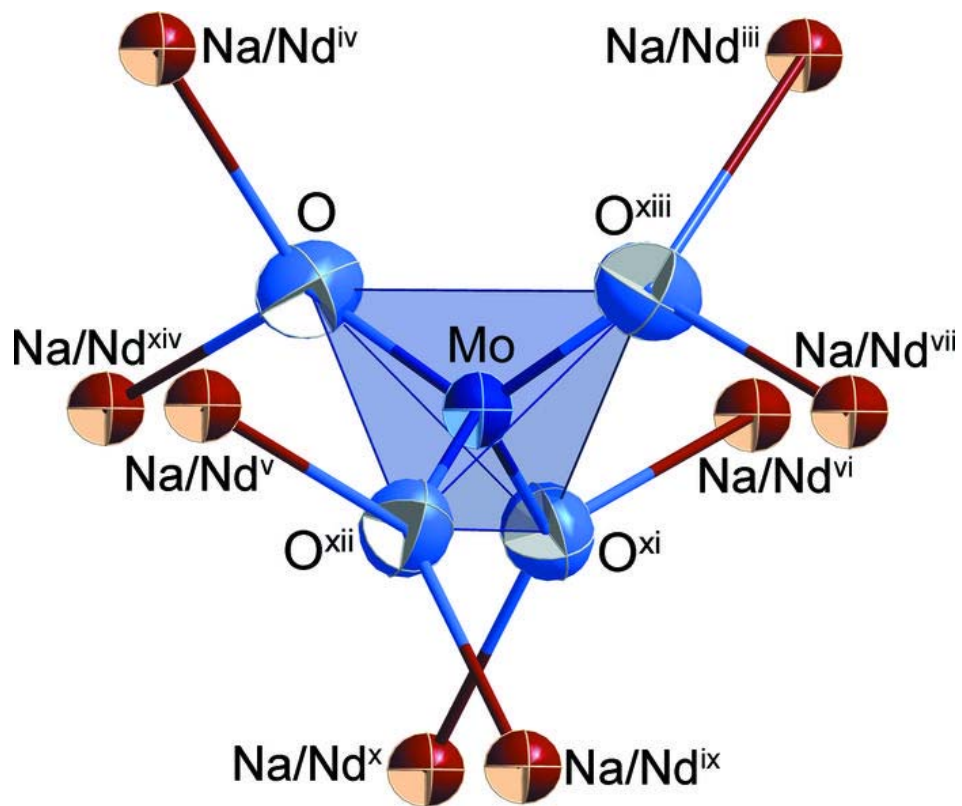


Fig. 3

